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10 When Vanadium(II) Compounds Are Dissolved In Water, The Following Equilibrium Is Established. $V^{2+} + H_2O \rightleftharpoons V^{3+} + H_2O + OH^-$ What Would Alter The Composition Of The Equilibrium Mixture In Favour Of The V^{2+} Ions? A Adding An Acid B Adding A Reagent That Selectively Precipitates
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Of Atomisation Of Silver = +285kJmol⁻¹ Enthalpy Change Of Atomisation Of Bromine = +112kJmol⁻¹ ... Apr 6th, 2024.

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