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3 Another Example A Solution With Two Volatile Components •Modified Raoult's Law • $VP_T = X_A P_A^* + X_B P_B^*$  Example •At 20.0 °C The Vapor Pressures Of Methanol (CH<sub>3</sub>OH)and Ethanol (C<sub>2</sub>H<sub>5</sub>OH) Are 95.0 And 45.0 mmHg Respectively. An Ideal Solution Contains 16.1 g Of Methanol Feb 6th, 2024

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B Raises Boiling Point Of Water From 100 °C To 100.64 °C •  $\Delta T_f$  Lowers Freezing Point Of Water From 0 °C To -2.32 °C •  $\pi$  Raises Osmotic Pressure From 1 Atm To 28.6 Atm While The Temperature Was Changing By A Degree Feb 7th, 2024

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If You Add 30.0g Of Salt To 3.75kg Of Water, What Will Be The Change In The Boiling Point Of Your Salt Water? Assume The  $K_b$  Of Water Is  $0.510^\circ\text{C}/m$ .  $30.03 \times .51$   
Assume You Have A  $3.60m$  Solution That Depressed The Freezing Point Of The Solution By  $0.8510^\circ\text{C}$ . What Is The Molal Freezing Poin May 4th, 2024

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Ions In Aqueous Solutions And Colligative Properties MIXED REVIEW SHORT ANSWER Answer The Following Questions In The Space Provided. 1. Match The Four Compounds On The Right To Their Descriptions On The Left. B An Ionic Compound That Is Quite Soluble In Water (a) HCl C An Ionic Compound That Is Not Very Soluble In Water (b)  $\text{NaNO}_3$  May 6th, 2024

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### **12.3 Colligative Properties - REMONDINI**

Colligative Properties For Molecular Solute Vs. Ionic Solute. Example: 1.0 M Of Sugar & 1.0 M Of Salt Show Different  $\Delta T_b$  And  $\Delta T_f$  Even Though The Molality Concentrations Of The Solute Are Identical. Note That;  $\Delta T = M \cdot K_{f,b}$  Where K Is Only A Function Of The Solvent !!! Reason For The Difference In Colligative Properties Must Be Based On ... Apr 6th, 2024

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