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Assigning Oxidation Numbers Directions: Assign Oxidation Numbers To All Of The Elements In Each Of The Compounds Or Ions Below. 1. HCl 11. H_2SO_3 2. KNO_3 12. H_2SO_4 3. OH^- 13. BaO 2 4. Mg_3N_2 14. KMnO_4 5. KClO_3 15. LiH 6. $\text{Al}(\text{NO}_3)_3$ 16. MnO_2 7. S 8 17. OF_2 8. H_2O_2 18. SO_3 9. PbO_2 19. NH_3 10. NaHSO_4 20. Na Chemistry IF8766 Page 91 ... May 1th, 2024

AP CHEMISTRY NOTES 2-1 ASSIGNING OXIDATION ...

ASSIGNING OXIDATION NUMBERS RULES FOR ASSIGNING OXIDATION NUMBERS: 1. The Oxidation Number Of Any Free Element (including Diatomic Elements) Is Always 0. 2. The Sum Of The Oxidation Number Of All Of The Atoms In A Compound Is Always Equal To 0. 3. The Sum Of The Oxidation Numbers Of All Of Th Apr 4th, 2024

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ASSIGNING OXIDATION NUMBERS WORKSHEET

4. The Sum Of The Oxidation Numbers Of All The Elements In A Polyatomic Ion Equals The Charge On The Ion. Directions: In The Following Questions, Give The Oxidation Number Of The Indicated Atoms/ion. 1. N In N_2O_3 2. S In H_2SO_4 3. C In $\text{C}_2\text{O}_4^{2-}$ 4. C In CO_2 5. Na In NaCl 6. H In H_2O Mar 4th, 2024

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2. For Monoatomic Ions, The Oxidation Number Is The Same As The Charge. 3. All Metals In Group 1 Form +1 Ions And Group 2 Form +2 Ions When Found In A Compound. 4. When Found In A Compound Fluorine Has A - 1 Oxidation State. 5. Hydrogen Has A +1 Oxidation State When Bonded To A Non Jul 4th, 2024

RULES FOR ASSIGNING OXIDATION NUMBERS

6. The Algebraic Sum Of The Oxidation Numbers Of Elements In A Compound Is Zero. CaF_2 H_2SO_4 H_2O +

$2(-1) = 0$ $2(+1) + S + 4(-2) = 0$ Find $S = +6$ For Ox.
No. Of S 7. The Algebraic Sum Of The Oxidation
Numbers Apr 4th, 2024

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Assigning Oxidation Numbers To Determine The Jan
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Assigning Oxidation Numbers ! KEY Determine The
Oxidation Numbers To ALL Of The Elements In Each Of
The Compounds Or Ions Below: 1) HCl 8) H₂O 2 15)
LiH H = +1 H = +1 (each) Li = +1 Cl = -1 O = -1
(each) H = -1 ... Microsoft Word - Assigning
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Numbe R In IMA Name Chemical Formula (unformatted) Coll.

Arseniosiderite $\text{A2Fe}^{+++}3(\text{AsO4})3\text{O2} \cdot 3(\text{H2O}) > 2$
Arsenohauchecornite Ni18Bi3AsS16 1 Arsenolamprite
As 1 Arsenolite $\text{As2O3} > 2$ Arsenopyrite $\text{FeAsS} > 10$
Arthurite $\text{UFe}^{+++}2(\text{AsO4}, \text{PO4}, \text{SO4})2(\text{O}, \text{OH})2 \cdot 4(\text{H2O})$ 1
Artinite $\text{Mg2}(\text{O3})(\text{OH})2 \cdot 3(\text{H2O}) > 2$ Apr 3th, 2024

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