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By Mixing 100.0 G Of Water, H₂O, And 100.0 G Of Ethanol, C₂H₅OH. Determine The
Mole Fractions Of Each Substance. 2. The Molality Of An Aqueous Solution Of Sugar
(C₁₂H₂₂O₁₁) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3.
Calculate The Molality Of 25.0 Grams Of 2th, 2024.

K Is The Molal Freezing Point Depression And M Molality Is ...As 1 Atm = 1.013 ×
10⁵ Pa, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$ Hence C
 $= \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$ As 1 M³ =
1000 L, This Corresponds To $C = 1.5 \times 10^{-4} \text{ mol/L}$, 2024 Richard L. Walker Papers An
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