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Value Of Equilibrium Constants. Answer: This Is Because Molar Concentration Of A Pine Solid Or Liquid Is Independent Of The Amount Present. 4th, 2024Chapter 7 Equilibrium - NCERT HelpClass XI Chapter 7 - Equilibrium Chemistry Page 2 Of 65 Website: Www.vidhyarjan.com Email: Contact@vidhyarjan.com Mobile: 9999 249717 Head Office: 1/3-H-A-2, Street # 6, East Azad Nagar, Delhi-110051 (One Km From 'Welcome' Metro Station) Hence, Kcfor The Equilibrium Is. Question 7.3: 4th, 2024(Chapter 7) (Equilibrium)(Chapter – 7) (Equilibrium) (Class – XI) 32 Question 7.43: The Ionization Constant Of HF, HCOOH And HCN At 298K Are  $6.8 \times 10$ –4, 1.8 $\times$  10-4 And 4.8  $\times$  10-9 Respectively. Calculate The Ionization Constants Of The Corresponding Conjugate Base. Answer 7.43: It Is Known That, Given, Ka Of HF =  $6.8 \times 10$ -4 Hence, Kb Of Its Conjugate ... 4th, 2024. Chemistry Notes Class 11 Chapter 7 Equilibrium Part-2Www.ncerthelp.com (Visit For All Ncert Solutions In Text And Videos, CBSE Syllabus, Note And Many More) Chemistry Notes Class 11 Chapter 7 Equilibrium Part-2 The Equilibrium Established Between The Unionised Molecules And The Ions In The Solution Of Weak Electrolytes Is Called Ionic Equilibrium. E.g., CH 3 COOH ⇔ CH 3 COO+ H-Electrolytes 3th, 2024

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