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Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol}$
 $122.55 \text{ g} \times 2 \text{ mol} = 245.10 \text{ g}$
 $\text{KClO}_4 \times 32.00 \text{ g} \times 1 \text{ mol} = 32.00 \text{ g}$
 O_2
 $\% \text{ Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
 $3\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
 Percent purity = 62.8%
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In A Camping Stove? 2. How Many Moles Of O_2
 Molecules Should Be Supplied To Burn 1 Mol Of CH_4 ...
 Was At One Time Considered For Use As A Rocket Fuel.
 ... How Many Formula Units Of Calcium Oxide (CaO)
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 Answers Calcium $\text{C}_2\text{O}_2 + 3 \text{Cr}(3)$ When The Potassium
 Chlorate (KClO_3) Is Heated, It Decomposes To Form
 Gases KCl And Oxygen (O_2)
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