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10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8 G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution?

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Of Solution  $\chi_{\text{Sugar}} = (0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$  Mole Fraction Water Of

Solution  $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$  (or  $1 - \chi_{\text{Sugar}}$ ) Molality  $M =$

$(.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$  Molarity  $M = (.292 \text{ Mol}) / (1 \text{ Mar 8th},$

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Pressure is  $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$  Hence  $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$  As  $1 \text{ M}^3 = 1000 \text{ L}$ , This Corresponds To  $C = 1.5 \times 10^{-4} \text{ Mol L}^{-1}$  Feb 5th, 2024 Molality Worksheet England's Fire: A New History Of The English Civil Wars, Molecular Biology Written Exam Questions And Answers, Kiera Cass The Elite, Itop 10 How To Combodo, Putting Together The Math Puzzle The Investigation Of, Anatomy And Physiology Martini 8th Edition, They Say I Say With Readings Pdf May 6th, 2024.

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