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Compounds Are Dissolved In Water, The Following
Equilibrium Is Established. $\text{V}^{2+} + \text{H}_2\text{O} \rightleftharpoons \text{V}^{3+} + \text{H}^+$
 $+ \text{OH}^-$ What Would Alter The Composition Of The
Equilibrium Mixture In Favour Of The V^{2+} Ions? A

Adding An Acid B Adding A Reagent That Selectively
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Chemical Equation Representing The Lattice Energy Of
AgBr. (ii) Use The Following Data To Calculate A Value
For The Lattice Energy Of AgBr(s). First Ionisation
Energy Of Silver = $+731\text{kJmol}^{-1}$ Electron Affinity Of
Bromine = 325kJmol^{-1} Enthalpy Change Of
Atomisation Of Silver = $+285\text{kJmol}^{-1}$ Enthalpy Change
Of Atomisation Of Bromine = $+112\text{kJmol}^{-1}$... Apr 3th,
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