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CAUTION UNDER NO CIRCUMSTANCES SHOULD FUEL OF A LOWER OCTANE RATING THAN THAT SPECIFIED BELOW, OR AUTOMOTIVE FUEL (REGARDLESS OF OCTANE) BE USED. I Grade (and Color): 100 Minimum Grade Aviation Fuel (green). 100 Low Lead Aviation Fuel (blue) Is Also Approved. Apr 7th, 2024

UNIT #11: Acids And Bases PH And POH Neutralization ...Hydrogen Ion Concentration, Shown As  $[H^+]$ , For Water =  $1.0 \times 10^{-7}M$  Hydroxide Ion Concentration, Shown As  $[OH^-]$ , For Water =  $1.0 \times 10^{-7}M$  Self-ionization Produces A Tiny Number Of Ions But Explains How Pure Water Can Behave As A Very Weak Electrolyte. D) Water Is The Usual Solvent For Acids And Bases; The Dissociation Of Acids Or Bases In An Feb 10th, 2024.

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Solution Is Created By Measuring  $3.60 \times 10^3$  Moles Of NaOH And  $5.95 \times 10^{-4}$  Moles Of HCl Into A Container And Then Water Is Added Until The Final Volume Is 1.00 L. What Is The PH Of This Solution? Since There Is Both Acid And Base We Will Assume A 1 Mole Acid:1 Mole Base Ratio Of Neutralization. Mar 8th, 2024.

PH And POH - Ms. Mogck's Classroom - Ms. Mogck's Classroom  
KEY Chemistry: PH And POH Calculations  
Part 1: Fill In The Missing Information In The Table Below.

PH	[ H <sup>+</sup> ]	POH	[ OH <sup>-</sup> ]	ACID Or BASE?
3.78	$1.66 \times 10^{-4}$ M	10.22	$6.03 \times 10^{-11}$ M	Acid
3.41	$3.89 \times 10^{-4}$ M	10.59	$2.57 \times 10^{-11}$ M	Acid
8.81	$1.55 \times 10^{-9}$ M	5.19	$6.46 \times 10^{-6}$ M	Base
8.69	$2.04 \times 10^{-9}$ M	5.31	$4.88 \times 10^{-6}$ M	Base
8.46	$3.47 \times 10^{-9}$ M	5.54	$2.88 \times 10^{-6}$ M	Base

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