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Chemical ...=  $K[D]_1 X[B]_1 Y K[D]_2 X[B]_2 Y$  Equation 5  
In This Equation K Cancels Out.  $[D]_1 = [D]_2$  Because  
We Are Starting Off With The Same Initial  
Concentration Of A In Each Trial. The Order Of Species  
D Which Is X Is Also Constant. Canceling Terms We  
Have Left:  $\text{Rate}_1 \text{Rate}_2 = K[B]_1 Y K[B]_2 Y$  Equation 6  
Consider The Following Chemical Reaction: 4th,  
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Web.gccaz.eduB. Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds

Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules  $KE \geq E_A$ ,

Activation Energy - Energy Barrier That Molecules Have To Surmount In Order To React. Energy Is Needed To

Break Reactant Bonds (endothermic Process). 2th, 2024AP\* Chemistry CHEMICAL KINETICS -

QuiaChemical Identity - What Exactly Is Reacting?

Usually Ions Of Opposite Charge React Very Rapidly.

Also, The More Bonds Between Reacting Atoms In A Molecule, The Slower The Reaction Rate. Why? More

Energy Is Required To Separate The Molecule Into Its "bits". Substances With Strong Bonds (larger Bond

Energies) Will React Much More Slowly. 4th, 2024.  
Chapter 14 Chemical Kinetics  
3 Chemical Kinetics  
Factors That Affect Reaction Rates • Physical State Of  
The Reactants In Order To React, Molecules Must  
Come In Contact With Each Other. If The Reaction Is  
Happening Between A Solid And A Liquid It Will React  
Only On The Surface. The More Homogeneous The  
Mixture Of Reactants, The Faster The Molecules Can  
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